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The Clausius-Clapeyron equation: $dP_i \text{ sat } P_i \text{ sat} = \frac{h_i \text{ vap}}{RT^2} dT$

RT 2 or $\ln P_i \text{ sat } 101 \text{ kPa} [] " \# \$ \% \& ' = () h_i \text{ vap } R 1 T (1 373$

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[K] * + , - . / so $P_{i sat} = 101 \text{ kPa} [] () \exp " \# h_{i vap} R 1 T " 1$
373 [K] \$ % & ' () * + , - .

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Chapter 6 2 If the fluid is in thermodynamic equilibrium any thermodynamic variable for a pure substance, like pure water, can be written in terms of any two other thermodynamic variables, i.e. $p=p(p,T)$ (6.1.1) where the functional relationship in depends on the substance.

Chapter 6 Thermodynamics and the Equations of Motion

Chapter 6: Solution Thermodynamics and Principles of Phase Equilibria In all the preceding chapters we have focused primarily on thermodynamic systems comprising pure substances. However, in all of nature, mixtures are ubiquitous.

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